Review Periodic Table Test

1. What are the 2 parts of an atom?
2. What are the 3 particles of an atom?
3. What are these particles charged and where are they found in the atom?
4. What is the majority of an atom?
5. How is an atom related to an element?
6. How is an element different from a compound?
7. Two elements are in the same “group” on the PT, what does this tell you?
8. Two elements are in the same “period” on the PT, what does this tell you?
9. How is the atomic number organized on the PT?
10. What subatomic particle determines the chemical properties of an atom?
11. How are the particles in #10 organized on the PT?
12. An atom has 4protons and 5 neutrons. How many electrons does it have? What is its atomic number? What is its atomic mass? What element is it?
13. Draw a picture of an element? Compound? Mixture?
14. What is the difference in a metal, nonmetal, and metalloid?
15. On the PT how do you distinguish between metals, nonmetals, and metalloids?
16. As you move from left to right on the periodic table how does the atomic number change?
17. As you move from left to right on the periodic table how does the atomic mass change?
18. How are the electrons in the outer shell related to the periodic table?

I CAN read the periodic table.

1. Al, Si, P, and S are called what?
2. What is the atomic number of Al ?
3. What is the atomic mass of S ?
4. Which element (Al, Si, P, or S) has the most

 parts needed to make it?

1. What do you call the numbers in the blocks

 13, 14, 15, and 16?

**Use your periodic table for questions 26-31.**

1. Use the periodic table to identify the

 symbol of Boron.

1. Use the periodic table to identify the

 atomic number of Boron.

1. Use the periodic table to identify the

 atomic mass of Nitrogen?

1. How many protons does an atom of

 Boron have?